# Identifying and classifying equilibria

How many critical points does

 $x' = \sin x$ 

have in  $[-4\pi, 4\pi]$ ? Mark them as stable or unstable.

A B b

### Identifying and classifying equilibria

How many critical points does

 $x' = \sin x$ 

have in  $[-4\pi, 4\pi]$ ? Mark them as stable or unstable.

- A. 5 critical points, 2 stable.
- B. 8 critical points, 3 stable.
- C. 8 critical points, 4 stable.
- D. 9 critical points, 4 stable.
- E. 9 critical points, 5 stable.

# Identifying and classifying equilibria

How many critical points does

 $x' = \sin x$ 

have in  $[-4\pi, 4\pi]$ ? Mark them as stable or unstable.

- A. 5 critical points, 2 stable.
- B. 8 critical points, 3 stable.
- C. 8 critical points, 4 stable.
- D. 9 critical points, 4 stable.
- E. 9 critical points, 5 stable.

How would you classify the equilibrium solution of the equation  $y' = (1 - y)^2$ ?

## Solving a first-order ODE

The equation 
$$y' - 1 = xy^2 + x + y^2$$

- A. is linear.
- B. is autonomous.
- C. is separable.
- D. does not have a unique solution for a given initial condition.

3 🕨 🖌 3

# Solving a first-order ODE

The equation 
$$y' - 1 = xy^2 + x + y^2$$

- A. is linear.
- B. is autonomous.
- C. is separable.
- D. does not have a unique solution for a given initial condition.

Find the general solution of this equation.

#### Chemical reactions

A second order chemical reaction involves the interaction (collision) of one molecule of a substance P with one molecule of a substance Q to produce one molecule of a new substance X. Let p and q denote the initial concentrations of P and Q respectively, and let x(t) denote the concentration of X at time t. The rate at which X is produced is proportional to the product of amount of P and Q remaining in the system. Write down the differential equation governing the system.

#### Chemical reactions

A second order chemical reaction involves the interaction (collision) of one molecule of a substance P with one molecule of a substance Q to produce one molecule of a new substance X. Let p and q denote the initial concentrations of P and Q respectively, and let x(t) denote the concentration of X at time t. The rate at which X is produced is proportional to the product of amount of P and Q remaining in the system. Write down the differential equation governing the system.

If x(0) = 0, find the limiting value of x(t) as  $t \to \infty$  without solving the differential equation.

(日) (周) (三) (三)

## Chemical reactions

A second order chemical reaction involves the interaction (collision) of one molecule of a substance P with one molecule of a substance Q to produce one molecule of a new substance X. Let p and q denote the initial concentrations of P and Q respectively, and let x(t) denote the concentration of X at time t. The rate at which X is produced is proportional to the product of amount of P and Q remaining in the system. Write down the differential equation governing the system.

If x(0) = 0, find the limiting value of x(t) as  $t \to \infty$  without solving the differential equation.

- (a) p
- (b) q
- (c) max(p,q)
- (d)  $\min(p,q)$
- (e) (p+q)/2